MOLARITY - CALCULATE MOLARITY FROM PERCENTAGE AND DENSITY -- Quiz

Question 1.

The density of a 56.0% by weight aqueous solution of methanol (CH₃OH) is 0.9046 g/cm³. What is the molarity of the compound? (Atomic weights: C = 12.01, H = 1.008, O = 16.00).

Question 2.

The density of a 64.0% by weight aqueous solution of glycerol (C₃H₈O₃) is 1.1663 g/cm³. What is the molarity of the compound? (Atomic weights: C = 12.01, H = 1.008, O = 16.00).

Question 3.

A solution prepared by dissolving 50.0 g of cesium chloride (CsCl) in 50.0 g of water has a density of 1.58 g/cm³. What is the molarity of the solution? (Atomic weights: Cs = 132.9, Cl = 35.45, H = 1.008, O = 16.00).

Question 1 = 15.8
Question 2 = 8.11
Question 3 = 4.69