RAULT’S LAW TOTAL VAPOR PRESSURE

Question 1.

What is the total pressure at 25°C of a solution of 10.0 g of CH₂Cl₂ and 1.00 g C₆H₁₂ at 25°C if the vapor pressures of the pure solvents are 435 and 166 mm Hg respectively at 25°C?

(Atomic weights: C = 12.01, H = 1.008, Cl = 35.45).

Question 2.

What is the total pressure at 25°C of a solution of 25 g of C₆H₁₄ and 50 g of C₆H₁₂ at 25°C if the vapor pressures of the pure solvents are 151 and 98 mm Hg respectively at 25°C?

(Atomic weights: C = 12.01, H = 1.008, Cl = 35.45).

Question 3.

What is the total pressure at 25°C of a solution of 2.90 moles of C₆H₁₄ and 5.94 moles of C₆H₁₂ at 25°C if the vapor pressures of the pure solvents are 151 and 98 mm Hg respectively at 25°C?

(Atomic weights: C = 12.01, H = 1.008, Cl = 35.45).

Question 1 = 410  Question 2 = 115

Question 3 = 115